

Article

Density and Refractive Index of Binary Ionic Liquid Mixtures with Common Cations/Anions, along with ANFIS Modelling

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Abstract: Ionic liquids have many interesting properties as they share the properties of molten salts as well as organic liquids, such as low volatility, thermal stability, electrical conductivity, non-flammability, and much more. Ionic liquids are known to be good solvents for many polar and nonpolar solutes. Combined with their special properties, ionic liquids are good replacements for the conventional toxic and volatile organic solvents. Each ionic liquid has different properties than others. In order to alter, tune, and enhance the properties of ionic liquids, sometimes, it is necessary to mix different ionic liquids to achieve the desired properties. However, using mixtures of ionic liquids in chemical processes requires reliable estimations of the mixtures' physical properties such as refractive index and density. The ionic liquids used in this work are 1-butyl-3-methylimidazolium thiocyanate ([BMIM][SCN]), 1-butyl-3-methylimidazolium tetrafluoroborate ([BMIM][BF₄]), 1-hexyl-3-methylimidazolium tetrafluoroborate ([HMIM][BF₄]), and 1-hexyl-3-methylimidazolium hexafluorophosphate ([HMIM][PF₆]). These ionic liquids were supplied by Io-li-tec and used as received. However, new measurements for the density and refractive index were taken for the pure ionic liquids to be used as reference. In the present work, the densities and refractive indices of four different binary mixtures of ionic liquids with common cations and/or anions have been measured at various compositions and room conditions. The accuracy of different empirical mixing rules for calculation of the mixtures refractive indices was also studied. It was found that the overall absolute average percentage deviation from the ideal solution in the calculation of the molar volume of the examined binary mixtures was 0.78%. Furthermore, all of the examined mixing rules for the calculation of the refractive indices of the mixtures were found to be accurate. However, the most accurate empirical formula was found to be Heller's relation, with an average percentage error of 0.24%. Furthermore, an artificial intelligence model, an adaptive neuro-fuzzy inference system (ANFIS), was developed to predict the density and refractive index of the different mixtures studied in this work as well as the published literature data. The predictions of the developed model were analyzed by various methods including both statistical and graphical approaches. The obtained results show that the developed model accurately predicts the density and refractive index with overall R^2 , RMSE, and AARD% values of 0.968, 7.274, 0.368% and 0.948, 7.32×10^{-3} and 0.319%, respectively, for the external validation dataset. Finally, a variance-based global sensitivity analysis was formed using extended the Fourier amplitude sensitivity test (EFAST). Our modelling showed that the ANFIS model outperforms the best available empirical models in the literature for predicting the refractive index of the different mixtures of ionic liquids.

Keywords: refractive index; density; ionic liquids; mixtures; ANFIS



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1. Introduction

Ionic liquids are salts in the liquid state. Due to the strong ionic bond between the cation and the anion, most ionic compounds exist in the form of crystal lattice (solid state) at room temperature. However, many asymmetric organic based ionic liquids appear in liquid

form at room temperature due to the weaker ionic bonding between the cation and anion. Ionic liquids have many interesting properties, as they share the properties of molten salts as well as organic liquids, such as low volatility, thermal stability, electrical conductivity, non-flammability, and much more. Ionic liquids are known to be good solvents for many polar and nonpolar solutes. Combined with their special properties, ionic liquids are good replacements for the conventional toxic and volatile organic solvents. Each ionic liquid has different properties to others. In order to alter, tune, and enhance the properties of ionic liquids, sometimes, it is necessary to mix different ionic liquids to achieve the desired properties.

However, using mixtures of ionic liquids in chemical processes requires the reliable estimations of the mixtures' physical properties, such as refractive index and density. Many general mixing rules for liquids have been proposed so far, which describe specific types of systems. This emphasizes the need to examine these mixing rules with respect to ionic liquid mixtures and to determine the best mixing rule for each physical property. Furthermore, assuming ideal behavior for the mixtures simplifies their complexity. Hence, the validity and accuracy of an ideal mixture assumption are uncertain and need to be examined. This will give a clearer picture in the design stage about the uncertainty of the calculations and in estimating the safety factor required for the design.

In addition to the existing empirical rules for the estimation of the different properties of the ionic liquid mixtures, such as refractive index, artificial intelligence (AI) models such as artificial neural networks (ANNs), regression models, fuzzy logic, support vector machine (SVM), and neural-based fuzzy interference system (ANFIS), can be used to predict the behavior of such complex systems. Different techniques have already been applied for the calculation of the various thermophysical properties of ionic liquid mixtures. Among various AI techniques, ANN alone or along with evolutionary algorithms such as genetic algorithms (GAs) and particle swarm optimization (PSO) has been used effectively to predict the properties of ionic liquid binary mixtures [1–4]. These techniques can predict responses with high accuracy, regardless of variation in input parameters. Very few studies have assessed the ability of the ANFIS model for this purpose [5]. ANFIS, however, is a potential soft computational modelling technique that combines the power of ANN with fuzzy logic [6–8]. ANFIS, as with the ANN, learns from training data with any complex relationships, then implements the solutions on a fuzzy inference system (FIS). ANFIS can use FIS to define hidden layers and to improve its predictive ability; thus, it eliminates the difficulty of defining the hidden layers that often exist in ANN; however, the number of membership functions (MFs) should be defined per input. The number of MFs depends on the number of the existing dataset. Their ability to overcome problems with experimental and deterministic models makes them ideal for complex chemical processes. However, determining the structure of ANN has become a major challenge in developing a highly accurate model [9]. Consequently, Hosoz et al. [10] pointed out that ANFIS, according to other studies, can predict more accurately than ANN. However, in other studies, it has been stated that the accuracy of these techniques depends on how they are implemented and their applications [11]. Considering all the aforementioned facts, in the present study, we have made an attempt to develop and examine an ANFIS model to predict the properties of ionic liquid binary mixtures. Our measured density and refractive index of binary ionic liquid mixtures as well as the available published data on these properties have been used in this work.

2. Properties of Binary Liquid Mixtures

2.1. Molar Volume

For ideal liquid mixtures, the molar volume is defined as:

$$v_{ideal\ mix} = \sum x_i v_i \quad (1)$$

where v_{mix} is the molar volume of the mixture, v_i is the individual molar volume of component i , and x_i is mole fraction of component i . Many liquids do not form ideal

mixtures and hence do not show volume additivity. The deviation from ideality can be calculated by introducing excess molar volume term to the previous equation. Hence, the equation is written as:

$$v_{real\ mix} = v_{ideal\ mix} + v^{excess} \quad (2)$$

2.2. Mixture Density

The density of a mixture can be defined as the overall mass of mixture occupying one unit volume. It can be calculated by:

$$\rho_{mix} = \frac{m_{total}}{V_{total}} \quad (3)$$

where ρ_{mix} is the mixture density, m_{total} is the total mass of mixture, and V_{total} is the total volume of mixture. For ideal liquid mixtures, using volume additivity, this equation can be written in terms of an individual component's densities as follows:

$$\frac{1}{\rho_{ideal\ mix}} = \sum \frac{x_i}{\rho_i} \quad (4)$$

where ρ_i is the density of component i and x_i is the mass fraction of component i . The average absolute relative difference percent (AARD%) is a good statistical parameter for analyzing the success of the proposed correlations for different properties such as density, which is defined as:

$$AARD\% = \frac{100 \sum \left| \frac{\rho_{ideal\ mix} - \rho_{mix, exp}}{\rho_{mix, exp}} \right|}{n} \quad (5)$$

2.3. Refractive Index

There are several mixing rules for the refractive index of mixtures in the literature. The most well-known correlations which have been used in this work are the following:

- Lorentz–Lorenz relation (L–L)

$$\frac{n_{mix}^2 - 1}{n_{mix}^2 + 2} = \sum \frac{n_i^2 - 1}{n_i^2 + 2} \phi_i \quad (6)$$

where n_{mix} is the refractive index of the mixture, n_i is the refractive index of pure component i , and ϕ_i is the volume fraction of component i in the mixture.

- Gladstone–Dale (G–D)

$$n_{mix} - 1 = \phi_1(n_1 - 1) + \phi_2(n_2 - 1) + \dots + \phi_n(n_n - 1) \quad (7)$$

- Arago–Biot (A–B)

$$n_{mix} = \phi_1 n_1 + \phi_2 n_2 + \dots + \phi_n n_n \quad (8)$$

- Weiner's relation (W)

$$\frac{n_{mix}^2 - n_1^2}{n_{mix}^2 + 2n_1^2} = \sum_{i \neq 1} \frac{n_i^2 - n_1^2}{n_i^2 + 2n_1^2} \phi_i \quad (9)$$

- Heller's relation (H)

$$\frac{n_{mix} - n_1}{n_1} = \frac{3}{2} \sum_{i \neq 1} \frac{m_i^2 - 1}{m_i^2 + 2} \phi_i \quad (10)$$

where $m_i = \frac{n_i}{n_1}$.

3. Materials and Experiments

The ionic liquids used in this work are 1-butyl-3-methylimidazolium thiocyanate ([BMIM][SCN]), 1-butyl-3-methylimidazolium tetrafluoroborate ([BMIM][BF₄]), 1-hexyl-3-methylimidazolium tetrafluoroborate ([HMIM][BF₄]), and 1-hexyl-3-methylimidazolium hexafluorophosphate ([HMIM][PF₆]). These ionic liquids were supplied by Io-li-tec and used as received. However, new measurements for the density and refractive index were taken for the pure ionic liquids to be used as reference. Table 1 shows the purity of each ionic liquid reported by Io-li-tec and their molecular weights.

Table 1. Purity of the ionic liquids used in this work.

Ionic Liquid	Molecular Weight	Purity
[BMIM][SCN]	197.30	≥98%
[BMIM][BF ₄]	226.03	≥99%
[HMIM][BF ₄]	254.08	≥99%
[HMIM][PF ₆]	312.24	≥99%

A pipette was used to transfer exact amounts of ionic liquids in a graduated cylinder to measure the total volume, each with uncertainty of 0.05 and 0.1 cm³, respectively. A mass balance of an uncertainty of 1×10^{-4} g was used to measure the mass of the samples. The refractive index was measured using a programmable digital KEM Refractometer (Kyoto Electronics Manufacturing Co., LTD., Model RA 620, Kyoto, Japan) connected to a water bath to maintain constant temperature. The uncertainty in the refractive index measurement is 1×10^{-5} , and the uncertainty in the temperature measurement is 0.1 K. The density was measured using an Anton Paar U-tube Densitometer (Model DMA 4500 M, Anton Paar GmbH, Graz, Austria) connected to a water bath. The uncertainty in the density measurement is 5×10^{-5} g cm⁻³, while the uncertainty in the temperature measurement was 0.1 K.

The density and refractive index of all samples were measured at 20 °C for the pure ionic liquids using an Anton Paar U-tube Densitometer and a Programmable Digital KEM Refractometer, respectively. The temperature was constantly maintained at 20 °C by connecting both the densitometer and the refractometer to a water bath set to a temperature of 20 °C. Binary mixtures of different ionic liquids were prepared from different volume compositions. The mass of each ionic liquid as well as the total mass of the mixture were measured before transfer to the graduated cylinder. A droplet was taken to be used in the refractometer to measure the refractive index. The remainder of the mixture was used to measure the mixture density using the density meter.

4. The Developed ANFIS Model and EFAST Sensitivity Analysis

Fuzzy logic is an appropriate technique for solving complex and nonlinear problems. ANNs have a strong ability to learn from existing data. Fuzzy logic theory is a powerful tool to deal with uncertainty. In general, fuzzy logic is an approach to calculate the degree of accuracy instead of using two-state true or false logic. Therefore, combining fuzzy logic theory with ANNs could provide great results for describing complex patterns, which is called ANFIS modelling. FIS assigns fuzzy rules and a membership function (MF) which ANN can optimize. Takagi–Sugeno and Mamdani are two well-known structures of FISs. Each FIS consists of a number of fuzzy rules, and each fuzzy rule has an antecedent and a consequent. In Mamdani-type FISs, both antecedents and consequents include fuzzy values; however, in Takagi–Sugeno-type FISs, the consequent of fuzzy rules are in the form of a fixed number (for zero type Takagi–Sugeno FISs) or a linear combination of inputs; this type of FISs are often used in ANFIS modelling. In this study, an ANFIS model was developed using the Takagi–Sugeno structure with the Neuro-Fuzzy Designer app in MATLAB 2017a software [9]. The ANFIS structure comprises four layers, which are fuzzy, product, consequent calculation, and output layer. The first layer includes antecedent

fuzzy values (also known as membership functions), the inputs enter these fuzzy values, and a membership grade is produced for each fuzzy value in each rule. In layer 2, the product of membership grades in a fuzzy rule is used as the firing strength of the rule in this research. Then, in layer 4 (in parallel with layers 1 and 2), the output of each rule is calculated using the input values. In the last layer, the output of the FIS is calculated by the output values and firing strengths of all rules. In this research, the weighted sum of outputs was used, where the weight of each rule output is the firing strength of the same rule. The unknowns of this modelling problem were the number of rules and fuzzy values (membership functions) of rules and consequent parameters. These were identified by using the experimental data.

In this study, various membership functions were investigated for their use in the rules' antecedents, such as the triangular-shaped membership function (trimf) and others, for which the details can be found elsewhere [12]. The result of this comparison is given in Section 6.

To estimate the density parameter using ANFIS, 1670 datasets were collected from the literature, which were applied for model training and cross validation. Cross validation partition defines a random partition on a dataset. This partition is used to define training and validation sets for validating a model. In this study, the early stopping method was used for preventing overfitting and the total dataset partitioned into training (85%) and holdout cross validation (15%) partitions. The holdout cross validation method creates a random nonstratified partition for holdout validation on n observations. Then, to determine the predictive power of the model, the measured data from this study (16 samples) were introduced to the model as external validation or a test dataset. The same procedure was followed for modelling the refractive index parameter, in which 149 data were collected and utilized from the published data for training and validation. Finally, 16 new measured data from this study were entered into the model for external validation. The hybrid-learning algorithm (as detailed in [13]) was used for parameter estimation, and the validity of the optimized ANFIS model was evaluated using experimental test data. The coefficient of determination (R^2), the root-mean-squared error (RMSE), and average absolute relative deviation (AARD%) were calculated to evaluate the model's performance [14,15].

$$R^2 = \frac{1 - \sum_{i=1}^n (Y_p - Y_e)^2}{\sum_{i=1}^n (Y_p - \bar{Y}_e)^2} \quad (11)$$

$$RMSE = \sqrt{\frac{1}{n} \sum_{i=1}^n (Y_e - Y_p)^2} \quad (12)$$

$$AARD\% = \frac{100}{n} \sum_{i=1}^n \frac{|Y_e - Y_p|}{Y_e} \quad (13)$$

where Y_e , Y_p , and \bar{Y}_e show the actual response, predicted response, and the average of the actual response, respectively, and n is the number of datasets.

For complex nonlinear models, such as the artificial intelligence models, global sensitivity analysis (GSA) techniques are able to provide appropriate information about parameter sensitivity. In the present study, the extended Fourier amplitude sensitivity test (EFAST), which is one of the efficient and well-known GSA methods, was applied for sensitivity analysis in MATLAB 2017 software. FAST is a variance-based global sensitivity analysis technique, which is based on conditional variances, indicating the individual or interaction effects of the uncertain inputs on the response parameter [16].

5. Experiments

5.1. Pure Ionic Liquids

Density and refractive index of the pure ionic liquids measured at 20 °C are shown in Table 2.

Table 2. Density and refractive index of the pure ionic liquids at 20 °C.

Ionic Liquid	Refractive Index	Density (g/cm ³)
[HMIM][BF ₄]	1.43940	1.14859, 1.14511 ^a
[HMIM][PF ₆]	1.42257	1.29544, 1.29145 ^a
[BMIM][SCN]	1.54543	1.07254, 1.06954 ^a
[BMIM][BF ₄]	1.43223	1.20447, 1.20085 ^a

^a Literature data at 298.15 K [17].

These values are used as references for the pure components in the analysis. Ionic liquid mixtures prepared for the experiments were labeled, and their measured masses and volumes at 20 °C are shown in Table 3.

Table 3. Measured volume and mass of prepared ionic liquid mixtures.

Mole Fraction (x_1)	Components		Mass (g)		Total
	1	2	1	2	
0.813			13.78308	3.88632	17.66940
0.620	[HMIM][BF ₄]	[HMIM][PF ₆]	10.33731	7.77264	18.10995
0.421			6.89154	11.65896	18.55050
0.214			3.44577	15.54528	18.99105
0.769			13.78308	3.21762	17.00070
0.555	[HMIM][BF ₄]	[BMIM][SCN]	10.33731	6.43524	16.77255
0.357			6.89154	9.65286	16.54440
0.172			3.44577	12.87048	16.31625
0.772			13.78308	3.61341	17.39649
0.560	[HMIM][BF ₄]	[BMIM][BF ₄]	10.33731	7.22682	17.56413
0.361			6.89154	10.84023	17.73177
0.175			3.44577	14.45364	17.89941
0.753			15.54528	3.21762	18.76290
0.534	[HMIM][PF ₆]	[BMIM][SCN]	11.65896	6.43524	18.09420
0.337			7.77264	9.65286	17.42550
0.160			3.88632	12.87048	16.75680

Table 3 shows the volume and mass of each component in the mixture. The mass of each component was used to calculate the mass fraction and mole fraction needed for the rest of calculations.

5.2. Mixtures of Ionic Liquids

5.2.1. Molar Volume and Density of Ionic Liquid Mixtures

Table 4 shows the measured mixture density, ideal mixture density, AARD%, measured mixture molar volume, ideal mixture molar volume, excess molar volume, and AARD% of the prepared ionic liquids.

Table 4. Measured and calculated volumetric data for the ionic liquid mixtures studied in this work.

Mixture	ρ_{mix} (g/cm ³)	$\rho_{ideal\ mix}$ (g/cm ³)	v_{mix} (cm ³ /mol)	$v_{ideal\ mix}$ (cm ³ /mol)	v^{excess} (cm ³ /mol)
[HMIM][BF ₄] + [HMIM][PF ₆]					
0.813	1.18620	1.17796	223.34763	224.90918	−1.56154
0.620	1.21185	1.20733	227.88141	228.73380	−0.85239
0.421	1.23965	1.2367	232.13618	232.69074	−0.55456
0.214	1.27477	1.26607	235.17064	236.78701	−1.61637
[HMIM][BF ₄] + [BMIM][SCN]					
0.769	1.13934	1.13338	211.48711	212.59928	−1.11218
0.555	1.11587	1.11817	205.05542	204.63350	0.42191
0.357	1.09203	1.10296	199.21668	197.24310	1.97359
0.172	1.10823	1.08775	186.84919	190.36790	−3.51871
[HMIM][BF ₄] + [BMIM][BF ₄]					
0.772	1.14710	1.159766	215.93122	213.57352	2.35771
0.560	1.15424	1.170942	209.43286	206.44638	2.98648
0.361	1.17505	1.182118	200.98041	199.77956	1.20086
0.175	1.18286	1.193294	195.23652	193.52985	1.70667
[HMIM][PF ₆] + [BMIM][SCN]					
0.753	1.25270	1.25086	226.61410	226.94751	−0.33341
0.534	1.20984	1.20628	213.78904	214.41969	−0.63065
0.337	1.16933	1.1617	201.87652	203.20263	−1.32611
0.160	1.13755	1.11712	189.63286	193.10083	−3.46797
AARD% = 0.78058%			AARD% = 0.78264%		

5.2.2. Refractive Index of Ionic Liquid Mixtures

Table 5 shows the experimental refractive indices of the ionic liquid mixtures and the calculated values using different empirical mixing rules.

Table 5. Experimental refractive indices of the ionic liquid mixtures and the calculated values using different empirical mixing rules.

Binary Mixtures	Refractive Index (n)					
	Experimental	L–L	G–D	A–B	W	H
[HMIM][BF ₄] + [HMIM][PF ₆]						
0.813	1.43831	1.43599	1.43604	1.43604	1.43603	1.43603
0.620	1.43745	1.43267	1.43267	1.43267	1.43267	1.43266
0.421	1.42883	1.42932	1.42931	1.42931	1.42930	1.42929
0.214	1.42408	1.42598	1.42594	1.42594	1.42594	1.42592
[HMIM][BF ₄] + [BMIM][SCN]						
0.769	1.46016	1.45955	1.46061	1.46061	1.46038	1.46033
0.555	1.46524	1.48088	1.48182	1.48182	1.48145	1.48125
0.357	1.50059	1.50209	1.50302	1.50302	1.50266	1.50217
0.172	1.52000	1.52364	1.52423	1.52423	1.52395	1.52309
[HMIM][BF ₄] + [BMIM][BF ₄]						
0.772	1.43918	1.43801	1.43798	1.43797	1.43797	1.43797
0.560	1.43767	1.43655	1.43654	1.43654	1.43654	1.43654
0.361	1.43475	1.43513	1.43511	1.43511	1.43511	1.43510
0.175	1.43288	1.43372	1.43367	1.43367	1.43367	1.43367
[HMIM][PF ₆] + [BMIM][SCN]						
0.753	1.45442	1.44636	1.44715	1.44715	1.44683	1.44676
0.534	1.47034	1.47048	1.47172	1.47172	1.47123	1.47094
0.337	1.50643	1.49504	1.49629	1.49629	1.49579	1.49512
0.160	1.51558	1.52005	1.52086	1.52086	1.52053	1.51930
	ARD (%)	0.24810	0.25711	0.25711	0.25190	0.24188

As we can see from Table 5, the refractive index of the ionic liquid mixtures can be confidently estimated using any of the discussed empirical formulae since the average relative error is around 0.25% for each correlation. However, Heller's relation (H) best estimates the refractive index of these ionic liquid mixtures with an average relative error of 0.242%.

6. ANFIS Modelling

The ANFIS models, developed to predict the density and refractive index of ionic liquid mixtures, are shown in Figures 1A and 2A. In this structure, we have employed the Takagi–Sugeno-type FIS with Gaussian curve membership functions. The ANFIS model has seven inputs of temperature (T), pressure (P), BMIM (mole fraction of BMIM ion), HMIM (mole fraction of HMIM ion), BF4 (mole fraction of BF4 ion), PF6 (mole fraction of PF6 ion), and SCN (mole fraction of SCN ion) for density and refractive index prediction. The grid partition technique is used to generate the rules for the model. The obtained results showed that the Gaussian membership function (gaussmf), compared to other membership functions, has less prediction errors in both models. Figures 1B and 2B illustrate the RMSE of the models' training and validation stages. During the training and validation process of the ANFIS model, the RMSEs were plotted versus the number of epochs. The RMSE reduces and then remains constant after some epochs. The RMSE values of the validation step for the density and refractive index models were equal to 8.730 and 7.45×10^{-3} , respectively. The regression curves plotted in Figures 1C and 2C display the model outputs, which show a very satisfactory performance. This comparison showed a good consistency between the predicted and actual responses of the ANFIS models.

We have used some statistical parameters as shown in Table 6 to examine the performance of the developed models.

Table 6. The performance of the statistical parameters for the proposed models.

Parameters	ANFIS (Density)			ANFIS (Refractive index)		
	Train	Validation	Test	Train	Validation	Test
R^2	0.992	0.985	0.968	0.988	0.972	0.948
RMSE	6.035	8.730	7.274	4.47×10^{-3}	7.45×10^{-3}	7.32×10^{-3}
AARD%	0.437	0.649	0.368	0.242	0.444	0.319

It is noteworthy that the ranges of error parameters for two responses (density and refractive index) are suitable at the training, validation, and test stages. The use of more data in the training and validation stages for density resulted in high predictive power, therefore the test data that were externally validated (not used in the training stage) resulted in predictions with higher accuracy. Moreover, a variance-based global analysis was performed using an extended EFAST method to determine which parameter had the greatest effect on the density and refractive index [18]. According to Figure 3A, the first three main factors affecting density are PF6, SCN, and pressure. In addition, according to Figure 3B, the most effective input parameters for the refractive index were SCN, pressure, and PF6. The coefficient values of different parameters are given in Figure 3.

Figures 4 and 5 show the three-dimensional plots of the developed models for the density and refractive index (n). As shown in Figure 4A,B, the density parameter often decreases with increasing temperature and increases with increasing pressure. In addition, an increase in parameters BMIM and SCN lead to a decrease in density (Figure 4B,E). BF4 has very little effect on density (Figure 4D). HMIM and PF6 have a positive correlation with density, as shown in Figure 4C,E. It is evident from Figure 5A that temperature has a negative effect on refractive index, and pressure has shown a non-uniform effect on it. However, the effect of the temperature on the refractive index is very small compared to the effect of pressure. In addition, an increase in parameters BMIM, BF4, and SCN leads to a decrease in the refractive index (Figure 5B–D). HMIM has very little effect on

the refractive index (Figure 5C). PF6 has a positive correlation with the refractive index, which is shown in Figure 5D. The model dataset contains 1670 density values over a temperature range of 220–472 K and a pressure range of 100–300,000 kPa collected from the NIST ILThermo database [19]. The model dataset contains 1670 density values over a temperature range of 220–472 K and a pressure range of 100–300,000 kPa collected from the NIST ILThermo database [19]. Our dataset for the refractive index which contains 149 values also has been collected from the NIST ILThermo database [19] in the range of 280–353 K for temperature, and the range of 81–116 kPa for pressure. This collected database is available in the Supplementary file in excel.

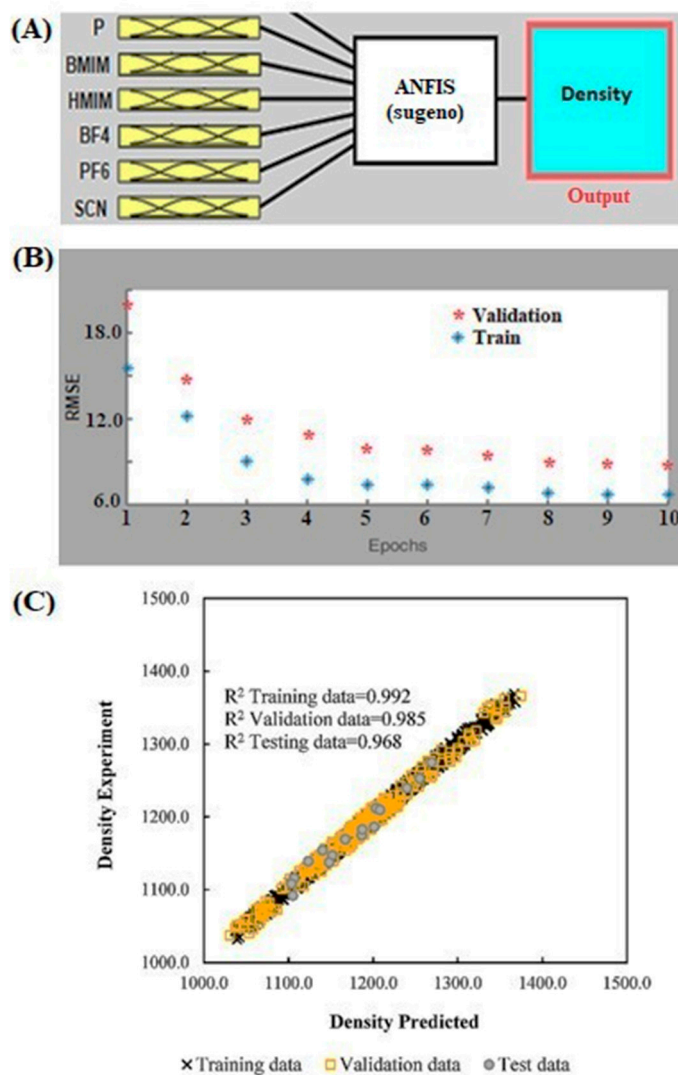


Figure 1. (A) Developed ANFIS structure, (B) training error curve, (C) regression between ANFIS predictions, and experimental results for density.

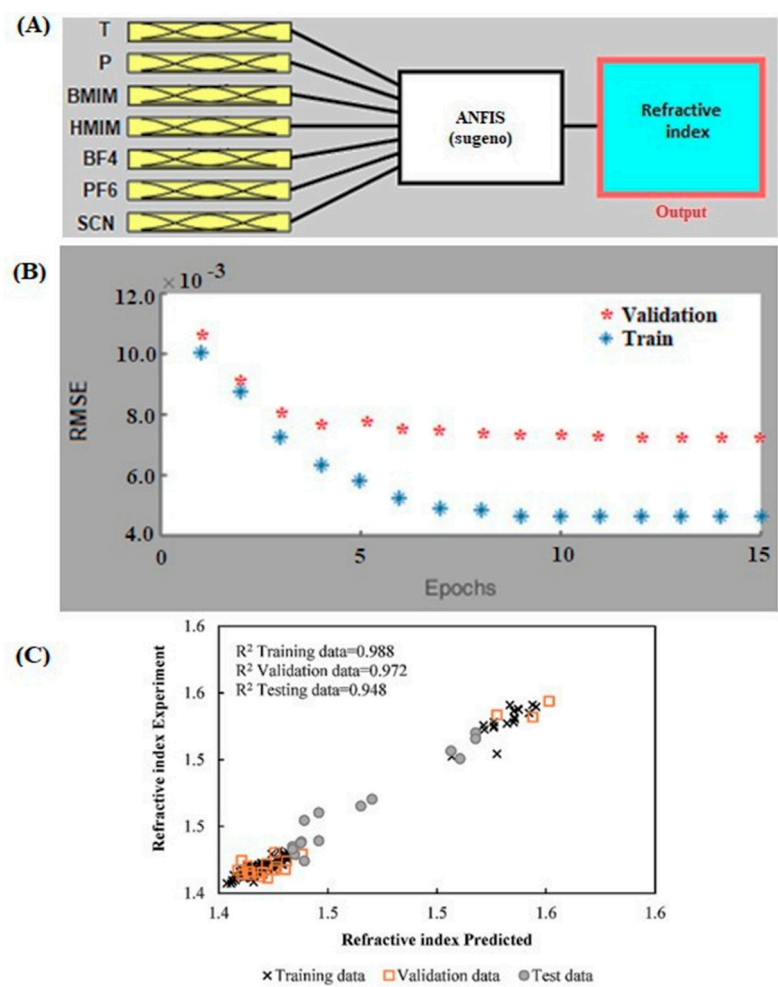


Figure 2. (A) Developed ANFIS structure, (B) training error curve, (C) regression between ANFIS predictions, and experimental results for refractive index.

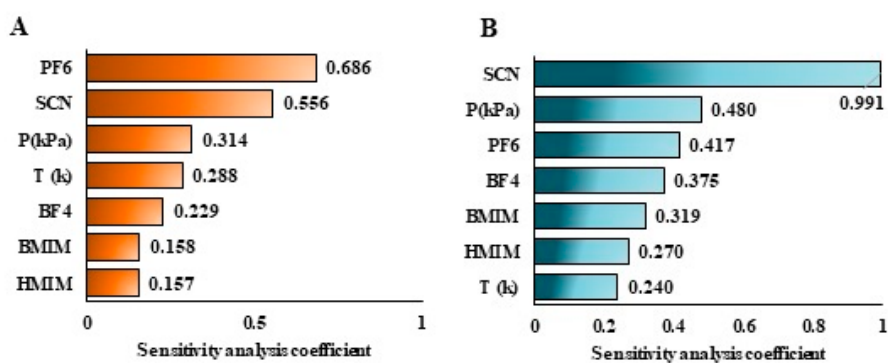


Figure 3. Relative importance of the input parameters on (A): density and (B): refractive index (n).

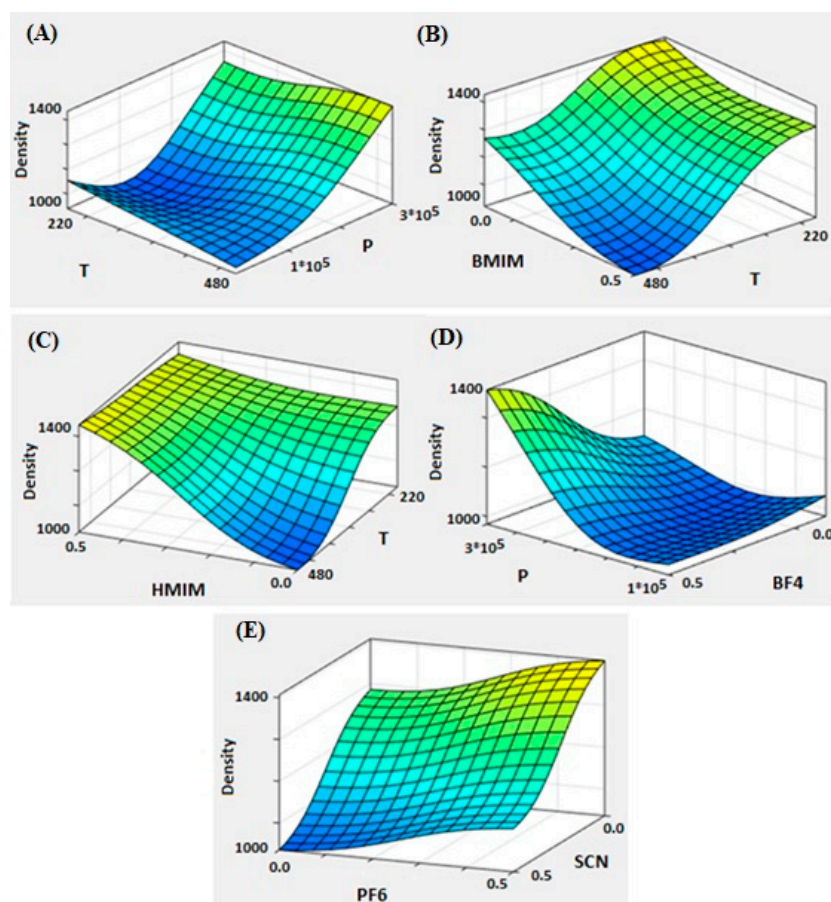


Figure 4. Three-dimensional mesh plot of the density variation using ANFIS model versus (A)–Temperature and Pressure, (B)–BMIM and Temperature, (C)–HMIM and Temperature, (D)–Pressure and BF4 and (E)–PF6 and SCN.

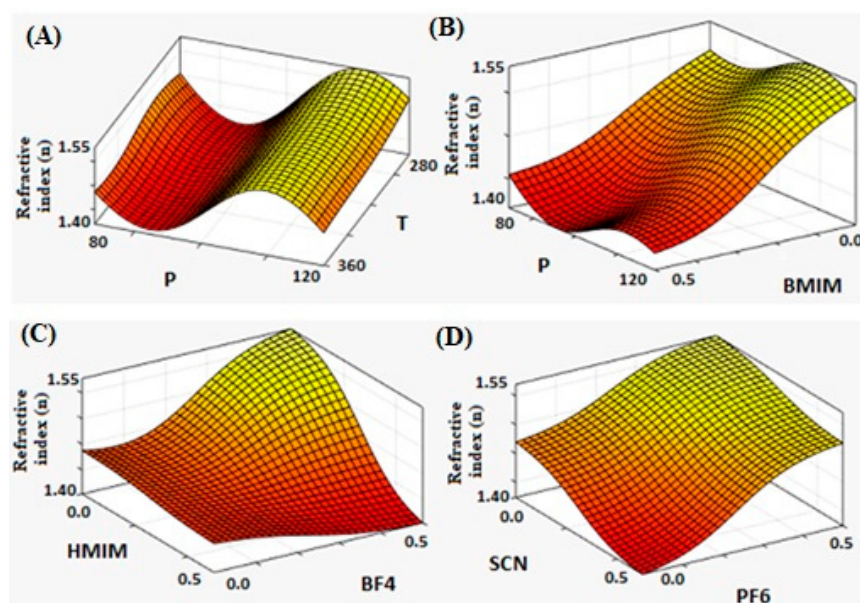


Figure 5. Three-dimensional mesh plot of the refractive index (n) variation using ANFIS model Versus, (A)–Pressure and Temperature, (B)–Pressure and BMIM, (C)–HMIM and BF4 and (D)–SCN and PF6.

7. Conclusions

Density and refractive index of different binary mixtures of ionic liquids were studied in this work. Some data were generated experimentally in this work and the rest of required data were collected from the literature. The effects of different variables such as temperature, pressure, and the molar concentrations of different functional groups in the ionic liquids were investigated. For analyzing the data, we first examined different models including empirical models for these properties. The deviation from the ideal mixture volume was calculated and found to be 0.78%. Therefore, it is a very accurate assumption to assume an ideal mixture and apply a volume additivity rule. Furthermore, the refractive indices of the binary ionic liquid mixtures were measured and compared with the estimated values using different empirical mixing rules. All of the empirical rules showed good performance for estimation of the refractive index of the binary ionic liquid mixtures, with an average percentage error around 0.25%. However, Heller's relation was found to be the best model to estimate these mixtures' refractive index with an average relative error of only 0.242%. For gaining a deep understanding of the effects of the different parameters on the density and refractive index of the studied mixtures, we have also developed artificial intelligence models. An ANFIS model was developed to predict a density and refractive index, and the grid partition technique was implemented. Our results showed that the Gaussian membership function, compared to other membership functions, has a low prediction error. Based on obtained results, the developed model showed excellent performance for predicting the density and refractive index of various mixtures, with overall R^2 , RMSE, and AARD% values of 0.985, 8.730, 0.649%, and 0.972, 7.45×10^{-3} , and 0.444%, in cross validation, and 0.968, 7.274, 0.368%, and 0.948, 7.32×10^{-3} , and 0.319% in the external validation stages, respectively. According to the sensitivity analysis using the EFAST method, the most effective input parameters for both properties were found to be SCN, pressure, and PF6.

Supplementary Materials: Supplementary data associated with this article can be found at: <https://www.mdpi.com/article/10.3390/liquids2040025/s1>.

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